

## Chemistry Stoichiometry Problem Sheet 2 Answer Key

Chemistry Stoichiometry Problem Sheet 2 Answer Key Chemistry Stoichiometry Problem Sheet 2 Answer Key This answer key provides detailed solutions and explanations for the problems presented in Chemistry Stoichiometry Problem Sheet 2 The problems focus on key concepts in stoichiometry including Mole conversions Converting between grams moles and particles atoms molecules or formula units Balancing chemical equations Ensuring the number of atoms of each element is equal on both sides of the equation Stoichiometric calculations Using mole ratios from balanced equations to determine the amount of reactants or products involved in a reaction Limiting reactants Identifying the reactant that limits the amount of product that can be formed Percent yield Calculating the efficiency of a reaction based on the actual yield obtained compared to the theoretical yield

**Problem 1 Mole Conversions**

a How many moles are in 250 g of sodium chloride NaCl Solution 1 Molar mass Find the molar mass of NaCl Na 2299 g/mol Cl 3545 g/mol 5844 g/mol 2 Conversion Use the molar mass to convert grams to moles 250 g NaCl 1 mol NaCl 5844 g NaCl 0427 mol NaCl

b How many grams are in 0750 mol of carbon dioxide CO<sub>2</sub> Solution 1 Molar mass Find the molar mass of CO<sub>2</sub> C 1201 g/mol 2O 1600 g/mol 4401 g/mol 2 Conversion Use the molar mass to convert moles to grams 2 0750 mol CO<sub>2</sub> 4401 g CO<sub>2</sub> 1 mol CO<sub>2</sub> 330 g CO<sub>2</sub>

c How many particles are in 150 mol of water H<sub>2</sub>O Solution 1 Avogadro's number Use Avogadro's number 6022 x 10<sup>23</sup> particles/mol to convert moles to particles 150 mol H<sub>2</sub>O 6022 x 10<sup>23</sup> molecules H<sub>2</sub>O 1 mol H<sub>2</sub>O 903 x 10<sup>23</sup> molecules H<sub>2</sub>O

**Problem 2 Balancing Chemical Equations**

Balance the following chemical equation CH<sub>4</sub> + O<sub>2</sub> → CO<sub>2</sub> + H<sub>2</sub>O Solution 1 Count atoms On the reactant side we have 1 C 4 H and 2 O On the product side we have 1 C 2 H and 3 O 2 Adjust coefficients To balance the equation we need to adjust the coefficients in front of each molecule CH<sub>4</sub> + 2O<sub>2</sub> → CO<sub>2</sub> + 2H<sub>2</sub>O 3 Recount atoms Now we have 1 C 4 H and 4 O on both sides of the equation

**Problem 3 Stoichiometric Calculations**

Consider the balanced equation 2Na + Cl<sub>2</sub> → 2NaCl

a How many grams of sodium chloride NaCl can be produced from 500 g of sodium Na Solution 1 Molar mass Find the molar masses of Na 2299 g/mol and NaCl 5844 g/mol 2 Convert grams to moles 500 g Na 1 mol Na 2299 g Na 0217 mol Na 3 Use mole ratio The balanced equation shows a 2:2 mole ratio between Na and NaCl This means 0217 mol Na will produce 0217 mol NaCl 4 Convert moles to grams 0217

mol NaCl 5844 g NaCl 1 mol NaCl 127 g NaCl b How many grams of chlorine gas  $\text{Cl}_2$  are required to react completely with 100 g of sodium Na Solution 1 Convert grams to moles 100 g Na 1 mol Na 2299 g Na 0435 mol Na 2 Use mole ratio The balanced equation shows a 2:1 mole ratio between Na and  $\text{Cl}_2$  This 3 means 0435 mol Na will react with 0217 mol  $\text{Cl}_2$  3 Convert moles to grams 0217 mol  $\text{Cl}_2$  7090 g  $\text{Cl}_2$  1 mol  $\text{Cl}_2$  154 g  $\text{Cl}_2$  Problem 4 Limiting Reactants Suppose 200 g of magnesium Mg reacts with 300 g of oxygen  $\text{O}_2$  according to the following balanced equation  $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$  a Which reactant is the limiting reactant Solution 1 Convert grams to moles 200 g Mg 1 mol Mg 2431 g Mg 0823 mol Mg 300 g  $\text{O}_2$  1 mol  $\text{O}_2$  3200 g  $\text{O}_2$  0938 mol  $\text{O}_2$  2 Determine limiting reactant Mg to MgO ratio 2 moles Mg 2 moles MgO 1  $\text{O}_2$  to MgO ratio 1 mole  $\text{O}_2$  2 moles MgO 12 Calculate moles of MgO produced 0823 mol Mg 2 mol MgO 2 mol Mg 0823 mol MgO 0938 mol  $\text{O}_2$  2 mol MgO 1 mol  $\text{O}_2$  1876 mol MgO The limiting reactant is Mg as it produces less MgO b How many grams of magnesium oxide MgO can be produced Solution 1 Use limiting reactant Since Mg is the limiting reactant we use its moles 0823 mol to calculate MgO production 2 Convert moles to grams 0823 mol MgO 4030 g MgO 1 mol MgO 332 g MgO Problem 5 Percent Yield In a reaction 150 g of sodium chloride NaCl is theoretically expected However only 125 g of NaCl is actually produced What is the percent yield Solution 1 Percent yield formula Percent yield =  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  2 Plug in values Percent yield =  $\frac{125 \text{ g}}{150 \text{ g}} \times 100 = 83.3$  Conclusion This answer key provides a comprehensive understanding of the concepts and calculations involved in stoichiometry By working through these problems and understanding the 4 explanations you can gain a strong foundation in solving stoichiometry problems in chemistry Remember to Always balance chemical equations before performing any stoichiometric calculations Use mole ratios from balanced equations to determine the amounts of reactants and products involved Identify the limiting reactant to determine the maximum amount of product that can be formed Calculate percent yield to assess the efficiency of a reaction Stoichiometry is a crucial aspect of chemistry enabling us to predict and analyze chemical reactions With practice and a solid understanding of the concepts you can master this important skill

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