

## Chapter 19 Acids Bases Study Guide Answer Key

Chapter 19 Acids Bases Study Guide Answer Key Chapter 19 Acids Bases and pH Study Guide Answer Key This study guide provides answers and explanations for key concepts covered in Chapter 19 on acids bases and pH. It's designed to help you reinforce your understanding and prepare for quizzes or exams.

I. What are acids and bases? Acids are substances that release hydrogen ions (H<sup>+</sup>) when dissolved in water. They typically have a sour taste and can react with metals to produce hydrogen gas. Bases are substances that release hydroxide ions (OH<sup>-</sup>) when dissolved in water. They generally have a bitter taste and feel slippery to the touch.

II. What is the pH scale? The pH scale is a logarithmic scale that measures the acidity or basicity of a solution. pH 0-6 indicates an acidic solution with a high concentration of H<sup>+</sup> ions. pH 7 is neutral with an equal concentration of H<sup>+</sup> and OH<sup>-</sup> ions. pH 8-14 indicates a basic solution with a high concentration of OH<sup>-</sup> ions.

III. Properties of Acids and Bases

- List the key properties of acids: Sour taste (Think of lemon juice or vinegar).
- React with metals: Producing hydrogen gas often seen as bubbles.
- Turn litmus paper red: Litmus paper is a pH indicator that changes color based on the acidity or alkalinity of the solution.
- React with carbonates: Producing carbon dioxide gas often seen as bubbles.
- Donate protons (H<sup>+</sup>): This is the Brønsted-Lowry definition of an acid.
- List the key properties of bases: Bitter taste (Like baking soda or soap).
- Slippery feel: Think of soap or laundry detergent.
- Turn litmus paper blue: Litmus paper again acts as a pH indicator.
- React with acids: Neutralizing each other and forming salt and water.
- Accept protons (H<sup>+</sup>): This is the Brønsted-Lowry definition of a base.

IV. Acid-Base Theories

- Describe the Arrhenius theory of acids and bases: This theory focuses on the production of ions in solution. Acids are defined as substances that produce H<sup>+</sup> ions in water. Bases are defined as substances that produce OH<sup>-</sup> ions in water.
- Describe the Brønsted-Lowry theory of acids and bases: This theory focuses on the transfer of protons (H<sup>+</sup>). Acids are defined as proton donors. Bases are defined as proton acceptors.
- What is a conjugate acid-base pair? A conjugate acid-base pair is formed when an acid loses a proton (H<sup>+</sup>) to form a conjugate base or when a base gains a proton to form a conjugate acid.
- For example, in the reaction of HCl (acid) and H<sub>2</sub>O (base), HCl donates a proton forming its conjugate base (Cl<sup>-</sup>). H<sub>2</sub>O accepts a proton forming its conjugate acid (H<sub>3</sub>O<sup>+</sup>).
- IV. pH and pOH

  - What is the relationship between pH and pOH? pH and pOH are inversely related, meaning that as one increases, the other decreases.
  - They are connected by the following equation: pH + pOH = 14.
  - How do you calculate pH and pOH? pH can be calculated using the following formula: pH = -log[H].
  - pOH can be calculated similarly: pOH = -log[OH].

is the concentration of hydroxide ions in moles per liter. What is a buffer solution? A buffer solution is a solution that resists changes in pH when small amounts of acid or base are added to it. Buffer solutions are important in biological systems maintaining a stable pH for enzymes and other biological processes.

3 V AcidBase Reactions

What is neutralization reaction? A neutralization reaction is a reaction between an acid and a base that produces salt and water. The products of neutralization are often neutral but this depends on the strengths of the acid and base involved.

What is a titration? Titration is a laboratory technique used to determine the concentration of an unknown solution by reacting it with a solution of known concentration. It involves carefully adding a known volume of a solution of known concentration (the titrant) to the unknown solution (the analyte) until the reaction is complete, typically indicated by a color change or pH measurement.

VI AcidBase Strength

What is the difference between a strong acid and a weak acid? Strong acids completely ionize dissociate in water meaning they donate all their protons. Examples include hydrochloric acid (HCl), nitric acid (HNO<sub>3</sub>) and sulfuric acid (H<sub>2</sub>SO<sub>4</sub>). Weak acids only partially ionize in water meaning they only donate a portion of their protons. Examples include acetic acid (CH<sub>3</sub>COOH), carbonic acid (H<sub>2</sub>CO<sub>3</sub>) and citric acid (C<sub>6</sub>H<sub>8</sub>O<sub>7</sub>).

What is the difference between a strong base and a weak base? Strong bases completely ionize in water meaning they accept all the protons they can. Examples include sodium hydroxide (NaOH) and potassium hydroxide (KOH). Weak bases only partially ionize in water meaning they only accept a portion of the available protons. Examples include ammonia (NH<sub>3</sub>), bicarbonate (HCO<sub>3</sub>) and carbonate (CO<sub>3</sub><sup>2-</sup>).

VII Applications of Acids and Bases

List some important applications of acids and bases.

Industrial processes: Acids and bases are used in many industrial processes such as manufacturing food production and metal refining.

Household uses: Acids and bases are common household products. For example, vinegar (acetic acid) is used for cleaning, baking soda (sodium bicarbonate) is used for baking and cleaning, and ammonia is used for cleaning.

Biological systems: Acids and bases play crucial roles in biological systems such as maintaining the pH of blood, digestion, and nerve function.

VIII Safety Precautions

Why is it important to handle acids and bases with caution? Acids and bases can be corrosive and can cause severe burns if they come into contact with skin or eyes. Some acids and bases can release harmful fumes or gases. It's crucial to wear appropriate protective gear when handling acids and bases such as gloves, goggles, and lab coats.

IX Review Questions

What is the difference between the Arrhenius and BrnstedLowry definitions of acids and bases? What are the properties of strong acids and weak acids? How do you calculate pH and pOH? What is the role of a buffer solution? Describe a neutralization reaction. How is it used in titration? Give examples of how acids and bases are used in our daily lives. Remember to consult your textbook and class notes for further details and examples. Good luck with your studies!

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Acid Base Regulation

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