

Chapter 19 Acids Bases Study Guide Answer Key

Chapter 19 Acids Bases Study Guide Answer Key Chapter 19 Acids Bases and pH Study Guide Answer Key This study guide provides answers and explanations for key concepts covered in Chapter 19 on acids bases and pH Its designed to help you reinforce your understanding and prepare for quizzes or exams I What are acids and bases Acids are substances that release hydrogen ions H when dissolved in water They typically have a sour taste and can react with metals to produce hydrogen gas Bases are substances that release hydroxide ions OH when dissolved in water They generally have a bitter taste and feel slippery to the touch What is the pH scale The pH scale is a logarithmic scale that measures the acidity or basicity of a solution pH 06 indicates an acidic solution with a high concentration of H ions pH 7 is neutral with an equal concentration of H and OH ions pH 814 indicates a basic solution with a high concentration of OH ions II Properties of Acids and Bases List the key properties of acids Sour taste Think of lemon juice or vinegar React with metals Producing hydrogen gas often seen as bubbles Turn litmus paper red Litmus paper is a pH indicator that changes color based on the acidity or alkalinity of the solution React with carbonates Producing carbon dioxide gas often seen as bubbles Donate protons H This is the BrnstedLowry definition of an acid List the key properties of bases Bitter taste Like baking soda or soap Slippery feel Think of soap or laundry detergent Turn litmus paper blue Litmus paper again acts as a pH indicator React with acids Neutralizing each other and forming salt and water 2 Accept protons H This is the BrnstedLowry definition of a base III AcidBase Theories Describe the Arrhenius theory of acids and bases This theory focuses on the production of ions in solution Acids are defined as substances that produce H ions in water Bases are defined as substances that produce OH ions in water Describe the BrnstedLowry theory of acids and bases This theory focuses on the transfer of protons H Acids are defined as proton donors Bases are defined as proton acceptors What is a conjugate acidbase pair A conjugate acidbase pair is formed when an acid loses a proton H to form a conjugate base or when a base gains a proton to form a conjugate acid For example in the reaction of HCl acid and H₂O base HCl H Cl HCl donates a proton forming its conjugate base Cl H₂O H H₃O H₂O accepts a proton forming its conjugate acid H₃O IV pH and pOH What is the relationship between pH and pOH pH and pOH are inversely related meaning that as one increases the other decreases They are connected by the following equation pH pOH 14 How do you calculate pH and pOH pH can be calculated using the following formula pH log₁₀H where H is the concentration of hydrogen ions in moles per liter

pOH can be calculated similarly $pOH = -\log[OH^-]$ where OH is the concentration of hydroxide ions in moles per liter. What is a buffer solution? A buffer solution is a solution that resists changes in pH when small amounts of acid or base are added to it. Buffer solutions are important in biological systems, maintaining a stable pH for enzymes and other biological processes.

3 V Acid-Base Reactions: What is a neutralization reaction? A neutralization reaction is a reaction between an acid and a base that produces salt and water. The products of neutralization are often neutral, but this depends on the strengths of the acid and base involved.

What is a titration? Titration is a laboratory technique used to determine the concentration of an unknown solution by reacting it with a solution of known concentration. It involves carefully adding a known volume of a solution of known concentration, the titrant, to the unknown solution, the analyte, until the reaction is complete, typically indicated by a color change or pH measurement.

VI Acid-Base Strength: What is the difference between a strong acid and a weak acid? Strong acids completely ionize in water, meaning they donate all their protons. Examples include hydrochloric acid (HCl), nitric acid (HNO₃), and sulfuric acid (H₂SO₄). Weak acids only partially ionize in water, meaning they only donate a portion of their protons. Examples include acetic acid (CH₃COOH), carbonic acid (H₂CO₃), and citric acid (C₆H₈O₇).

What is the difference between a strong base and a weak base? Strong bases completely ionize in water, meaning they accept all the protons they can. Examples include sodium hydroxide (NaOH) and potassium hydroxide (KOH). Weak bases only partially ionize in water, meaning they only accept a portion of the available protons. Examples include ammonia (NH₃), bicarbonate (HCO₃⁻), and carbonate (CO₃²⁻).

VII Applications of Acids and Bases: List some important applications of acids and bases. Industrial processes: Acids and bases are used in many industrial processes, such as manufacturing, food production, and metal refining. Household uses: Acids and bases are common household products. For example, vinegar (acetic acid) is used for cleaning, baking soda (sodium bicarbonate) is used for baking and cleaning, and ammonia is used for cleaning.

Biological systems: Acids and bases play crucial roles in biological systems, such as maintaining the pH of blood, digestion, and nerve function.

VIII Safety Precautions: Why is it important to handle acids and bases with caution? Acids and bases can be corrosive and can cause severe burns if they come into contact with skin or eyes. Some acids and bases can release harmful fumes or gases. It's crucial to wear appropriate protective gear when handling acids and bases, such as gloves, goggles, and lab coats.

IX Review Questions: What is the difference between the Arrhenius and Brønsted-Lowry definitions of acids and bases? What are the properties of strong acids and weak acids? How do you calculate pH and pOH? What is the role of a buffer solution? Describe a neutralization reaction. How is it used in titration? Give examples of how acids and bases are used in our daily lives. Remember to consult your textbook and class notes for further details and examples. Good luck with your studies!

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all persons eligible for covid 19 vaccination may book an appointment at any participating general practitioner gp clinic through vaccine gov sg covid or at a polyclinic through the healthhub

may 13 2025 individuals at increased risk of severe covid 19 such as those aged 60 years and above medically vulnerable individuals or residents of aged care facilities are recommended to keep

oct 25 2024 to bring covid 19 vaccination closer to the community close to 500 healthier sg gp clinics and 10 polyclinics located island wide are now providing covid 19 vaccination services

may 18 2024 since covid 19 vaccination started in 2020 2021 the vaccines have consistently been proven to be safe and effective in protecting individuals from severe illness

oct 28 2023 the expert committee on covid 19 vaccination ec19v recommends that the pfizer biontech comirnaty and moderna spikevax vaccines used for covid 19 vaccination in singapore

apr 3 2020 the taskforce has reviewed the latest evidence and pattern of transmission of covid 19 in singapore most of the cases have arisen from persons coming into close contact with an infected

dec 3 2024 novavax is in the process of filing for regulatory approval for its updated covid

19 vaccine formulation further updates will be provided when novavax has successfully filed for

jan 23 2025 from 24 january to 30 april 2025 the updated jn 1 novavax nuvaxovid covid 19 vaccine will be available at selected healthier sg general practitioner gp clinics under the

dec 15 2023 as covid 19 testing is no longer routinely required in dorscon green the number of reported covid 19 cases is expected to under represent the actual situation

oct 15 2022 vaccination is our primary defence against covid 19 and given that immunity in the population is likely waning over time we urge the public to get vaccinated or boosted when they are

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