

## Chapter 13 States Of Matter Quiz

Chapter 13 States Of Matter Quiz Chapter 13 States of Matter Quiz  
Instructions Choose the best answer for each question You may use your notes and textbook Multiple Choice 1 point each 1 What is the state of matter that has a definite shape and volume a Solid b Liquid c Gas d Plasma 2 Which state of matter is characterized by particles that are closely packed and vibrating in fixed positions a Solid b Liquid c Gas d Plasma 3 What is the process called when a liquid changes into a gas a Freezing b Melting c Vaporization d Condensation 4 Which of the following properties is NOT a characteristic of a liquid a Definite volume b Definite shape c Ability to flow d Takes the shape of its container 5 What is the state of matter that exists at extremely high temperatures and is composed of ionized particles 2 a Solid b Liquid c Gas d Plasma 6 Which of the following is an example of a physical change a Burning wood b Baking a cake c Melting ice d Rusting iron 7 What is the process called when a gas changes directly into a solid a Sublimation b Deposition c Freezing d Condensation 8 Which of the following factors affects the rate of evaporation a Temperature b Surface area c Wind speed d All of the above 9 What is the term used to describe the pressure exerted by the atmosphere a Air pressure b Barometric pressure c Atmospheric pressure d All of the above 10 What is the relationship between temperature and the kinetic energy of particles a As temperature increases kinetic energy decreases b As temperature decreases kinetic energy increases c As temperature increases kinetic energy increases d Temperature and kinetic energy are unrelated True or False 1 point each 11 Solids are incompressible TrueFalse 3 12 Liquids have a definite shape TrueFalse 13 The particles in a gas are constantly moving and colliding TrueFalse 14 Plasma is the most common state of matter in the universe TrueFalse 15 The boiling point of a liquid is the temperature at which it changes into a solid TrueFalse Short Answer 2 points each 16 Explain the difference between a physical change and a chemical change 17 Describe the properties of a gas 18 Explain how pressure affects the state of matter 19 Give an example of sublimation and deposition in everyday life 20 Describe the role of intermolecular forces in the states of matter Bonus 3 points 21 Explain the relationship between the states of matter and the kinetic molecular theory Answer Key Multiple Choice 1 a Solid 2 a Solid 3 c Vaporization 4 b Definite shape 5 d Plasma 6 c Melting ice 7 b Deposition 8 d All of the above 9 d All of the above 10 c As temperature increases kinetic energy increases True or False 11 True 12 False 13 True 14 True 4 15 False Short Answer 16 A physical change is a change in the form or appearance of a substance but not its chemical composition A chemical change results in the formation of a new substance with different properties 17 Gases have no definite shape or volume they are highly compressible and their particles are far apart and move

randomly 18 Increased pressure can force particles closer together favoring a more condensed state like a liquid or solid Decreased pressure allows particles to spread further apart favoring a gaseous state 19 Sublimation Dry ice solid carbon dioxide sublimates into a gas at room temperature Deposition Frost forming on a window is an example of water vapor gas depositing as ice crystals solid 20 Intermolecular forces are attractive forces between molecules Stronger intermolecular forces lead to higher melting and boiling points as more energy is required to overcome these forces and change the state of matter Bonus 21 The kinetic molecular theory states that matter is made up of particles in constant motion The energy of these particles kinetic energy determines the state of matter In solids particles have low kinetic energy and are tightly packed In liquids particles have higher kinetic energy and are less tightly packed In gases particles have the highest kinetic energy and are far apart Grading Rubric Multiple Choice 1 point for each correct answer True or False 1 point for each correct answer Short Answer 2 points for each answer that demonstrates understanding of the concept and provides a clear explanation Bonus 3 points for a complete and accurate explanation of the relationship between the states of matter and the kinetic molecular theory Total Points 25 points 3 points bonus 5

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