

Thermodynamics And The Kinetic Theory Of Gases

Volume 3 Of Pauli Lectures On Physics Dover

Books On Physics

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this book can be described as a student's edition of the author's dynamical theory of gases it is written however with the needs of the student of physics and physical chemistry in mind and

those parts of which the interest was mainly mathematical have been discarded this does not mean that the book contains no serious mathematical discussion the discussion in particular of the distribution law is quite detailed but in the main the mathematics is concerned with the discussion of particular phenomena rather than with the discussion of fundamentals

kinetic theory volume i the nature of gases and of heat covers the developments in area of kinetic theory statistical mechanics and thermodynamics this book is organized into two parts encompassing 11 chapters the book starts with an overview of the history of atomism the caloric theory the conservation of energy the virial theorem and atomic magnitudes the second part deals first with the delineation of observed phenomena of motions through the repulsion theory this part also considers other forces of nature including fire and heat with emphasis on the nature of motion of these forces and their conservation this book is intended for physicists and physics teachers and students

appendices after each chapter

a pioneering text in its field this comprehensive study is one of the most valuable texts and references available the author explores the classical kinetic theory in the first four chapters with discussions of the mechanical picture of a perfect gas the mean free path and the distribution of molecular velocities the fifth chapter deals with the more accurate equations of state or van der waals equation and later chapters examine viscosity heat conduction surface phenomena and brownian movements the text surveys the application of quantum theory to the problem of specific heats and the contributions of kinetic theory to knowledge of electrical and magnetic properties of molecules concluding with applications of the kinetic theory to the conduction of electricity in gases 1934 edition

this monograph and text was designed for first year students of physical chemistry who require further details of kinetic theory the treatment focuses chiefly on the molecular basis of important thermodynamic properties of gases including pressure temperature and thermal energy includes numerous exercises many partially worked out and end of chapter problems 1966 edition

this book introduces physics students and teachers to the historical development of the kinetic theory of gases by providing a collection of the most important contributions by clausius maxwell and boltzmann with introductory surveys explaining their significance in addition extracts from the works of boyle newton mayer joule helmholtz kelvin and others show the historical context of ideas about gases energy and irreversibility in addition to five thematic essays connecting the classical kinetic theory with 20th century topics such as indeterminism and interatomic forces there is an extensive international bibliography of historical commentaries on kinetic theory thermodynamics etc published in the past four decades the book will be useful to historians of science who need primary and secondary sources to be conveniently available for their own research and interpretation along with the bibliography which makes it easier to learn what other

historians have already done on this subject

excerpt from a treatise on the kinetic theory of gases we are obliged therefore to abandon the strictly kinetic method and to adopt the statistical method about the publisher forgotten books publishes hundreds of thousands of rare and classic books find more at forgottenbooks com this book is a reproduction of an important historical work forgotten books uses state of the art technology to digitally reconstruct the work preserving the original format whilst repairing imperfections present in the aged copy in rare cases an imperfection in the original such as a blemish or missing page may be replicated in our edition we do however repair the vast majority of imperfections successfully any imperfections that remain are intentionally left to preserve the state of such historical works

this historic book may have numerous typos and missing text purchasers can usually download a free scanned copy of the original book without typos from the publisher not indexed not illustrated 1920 edition excerpt of gas from which the molecule comes and thus increases with decrease of pressure the difference between v_a and v_c therefore also increases with decrease of pressure which increases the value of the ratio t_j / i_j according to the foregoing equation this corresponds to a decrease of and thus the effect of slipping becomes the greater the lower the pressure the dynamical mechanism underlying the result that the viscosity of a gas is independent of its density may be gaseous viscosity and the temperature 127 illustrated by the following considerations a molecule transfers a certain amount of momentum to the gas at the end of each transfer distance on migrating at right angles to the motion of the gas in the direction of the decrease of motion if the concentration of the gas is halved the length of each transfer distance is doubled while the momentum transferred at the end of each transfer distance is also doubled since the velocity gradient of the gas remains the same since a change in molecular concentration of a gas does not alter the molecular velocities the momentum transferred per second by a molecule moving between two parallel plates of material one of which is at rest while the other moves parallel to itself is in the latter case double that in the former but since the number of molecules per cubic cm available for momentum transference in the former case is half the number in the latter the total momentum transferred is in each case the same or the viscosity of the gas has not been altered by altering its density the quantity k is a function of the temperature this is shown by the calculated values of k contained in table x for a number of gases at different temperatures these values are not corrected for

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