

# The Mole Webquest Answer Key

mole can be defined as the amount of substance that contains the same number of entities atoms molecules ions etc as the number of atoms present in 12g of the carbon 12 isotope

define mole fraction a solution of sucrose in water is labelled as 20 w/w what would be the mole fraction of each component in the solution

two components a and b form an ideal solution the mole fractions of a and b in ideal solution are  $x_a$  and  $x_b$  while that of in vapour phase these components have their mole fractions as

a triatomic linear neglect vibration degree of freedom gas of two mole is taken through a reversible process ideal starting from a as shown in the figure the volume ratio  $\frac{V_2}{V_1}$  v

consider the reaction  $2\text{SO}_2 + \text{O}_2 \rightleftharpoons 2\text{SO}_3$  for which  $K_c = 278\text{M}^{-1}$ . If 1.00 mole of each of the reagents  $\text{SO}_2$ ,  $\text{O}_2$  and  $\text{SO}_3$  are mixed in a 1.0 l flask determine the reaction

find the average translational kinetic energy per molecule if one mole of the gas is contained in a volume 1 23 10 3 m<sup>3</sup> at a pressure 2 105n m<sup>2</sup> avagadro s number is 6 02 1023

0.5 mole of  $\text{H}_2$  g and 1.0 mole of  $\text{H}_2\text{g}$  but no  $\text{I}_2$  are added to a 10 l vessel and allowed to reach equilibrium according to the following reaction  $\text{H}_2 + \text{I}_2 \rightleftharpoons 2\text{H}_2\text{g}$  if x is the equilibrium

the compressibility factor for 1 mole of a vander waals gas at 273 k 8 21 atm pressure having volume 1 092 litre is 0 4 assuming the volume of gas molecules to be negligible

the vapour pressure of two volatile liquid mixtures is  $P = 5 + 3x_B$  in mm of hg where  $x_B$  is mole fraction of B in mixture what is the ratio of mole fractions of A & B in liquid phase  $P_A/P_B$

an aqueous solution is prepared by dissolving 0.1 mole  $\text{H}_2\text{CO}_3$  in sufficient water to get 100 mL solution at 25°C for  $\text{H}_2\text{CO}_3 \text{Ka} = 4.0 \times 10^{-6}$  and  $\text{Ka} = 2.5 \times 10^{-11}$

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