

# Edexcel Igcse Chemistry Chapter 6 Page 51

## Answers

Edexcel Igcse Chemistry Chapter 6 Page 51 Answers Edexcel IGCSE Chemistry Chapter 6 Page 51 Answers and Explanations This article will provide answers and explanations for the questions found in Edexcel IGCSE Chemistry Chapter 6 page 51 The chapter focuses on the reactivity of metals a crucial concept in understanding chemical reactions and their applications Before we begin remember Understanding the concepts Relying solely on answers wont help you master the material Read the chapter carefully make notes and try to answer the questions yourself before consulting this guide Active learning Use this guide to reinforce your understanding Refer back to the chapter for relevant explanations and examples Practice makes perfect Solve additional practice problems to further solidify your knowledge Lets delve into the answers Question 1 a What is meant by the term reactivity series The reactivity series is a list of metals arranged in order of their decreasing reactivity It helps us understand how readily metals react with other substances particularly acids and water b How can the reactivity series be used to predict the outcome of a reaction between a metal and an acid The reactivity series helps predict the outcome of a reaction between a metal and an acid based on their relative positions Metals above hydrogen in the series These metals are more reactive than hydrogen They will react with acids to produce hydrogen gas and a salt Metals below hydrogen in the series These metals are less reactive than hydrogen They will not react with dilute acids c Why is the reactivity series useful for predicting the outcome of a reaction between a 2 metal and a solution of a metal salt The reactivity series also predicts the outcome of a reaction between a metal and a solution of a metal salt Heres how A more reactive metal will displace a less reactive metal from its salt solution For example iron Fe is more reactive than copper Cu If we add iron to a solution of copper sulfate  $\text{CuSO}_4$  the iron will displace copper from the solution forming iron sulfate  $\text{FeSO}_4$  and copper metal A less reactive metal will not displace a more reactive metal from its salt solution Conversely if we add copper to a solution of iron sulfate no reaction will occur Question 2 a Write a balanced symbol equation for the reaction between magnesium and dilute hydrochloric acid The balanced symbol equation for the reaction is  $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$  b Describe how you would carry out an experiment to investigate the reaction between magnesium and dilute hydrochloric acid Heres a detailed description of the experiment 1 Safety Precautions Wear safety goggles and gloves Ensure the experiment is carried out in a wellventilated area 2 Materials Magnesium ribbon cleaned with sandpaper Dilute hydrochloric acid HCl Test tube Delivery tube Beaker Bunsen burner Splint for testing the gas 3 Procedure Fill a test tube with dilute hydrochloric acid Carefully add a small piece of magnesium ribbon to the acid Observe the reaction

noting any changes in appearance and the production of gas bubbles Use the delivery tube to collect the gas produced in a beaker 3 Test the gas collected with a lit splint If the splint extinguishes with a squeaky pop it indicates the presence of hydrogen gas c What observations would you expect to make during this experiment During the experiment you would observe Fizzing The magnesium ribbon will react with the acid producing bubbles of hydrogen gas Dissolving The magnesium ribbon will gradually dissolve in the acid Temperature change The solution will become warm due to the exothermic reaction Gas test When the collected gas is tested with a lit splint the splint will extinguish with a squeaky pop confirming the presence of hydrogen gas Question 3 a Explain why the reaction between calcium and water is more vigorous than the reaction between magnesium and water Calcium is more reactive than magnesium This means that calcium has a greater tendency to lose electrons and form positive ions Consequently calcium reacts more vigorously with water releasing more energy in the form of heat and producing hydrogen gas more rapidly b Write a balanced symbol equation for the reaction between calcium and water The balanced symbol equation for the reaction is  $\text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$  c Why is it dangerous to add potassium to water Potassium is a highly reactive metal The reaction between potassium and water is extremely vigorous and releases a large amount of heat The heat can ignite the hydrogen gas produced leading to an explosion Question 4 a Explain how the reactivity series can be used to predict the outcome of a reaction between a metal and a solution of a metal salt As explained earlier a more reactive metal will displace a less reactive metal from its salt solution This is based on the tendency of a more reactive metal to lose electrons and form positive ions The more reactive metal will displace the less reactive metal from its ionic compound b Write a balanced symbol equation for the reaction between iron and copper sulfate 4 solution The balanced symbol equation for the reaction is  $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$  c Describe what you would observe if you added a piece of iron to a solution of copper sulfate You would observe Color change The blue color of the copper sulfate solution will fade as the copper ions are replaced by iron ions which form a pale green solution Formation of a reddishbrown solid A reddishbrown solid copper metal will start to deposit on the surface of the iron piece Question 5 a Why are metals used in the construction of buildings Metals are used in construction for various reasons Strength and durability They are strong and durable materials that can withstand high loads and resist wear and tear Corrosion resistance Some metals like stainless steel are resistant to rust and corrosion making them suitable for outdoor applications Malleability and ductility Metals are malleable can be hammered into shapes and ductile can be drawn into wires which allows for flexibility in design Conductivity Metals conduct heat and electricity well making them suitable for various electrical and heating applications b Why are some metals used in the manufacture of jewelry Metals are used in jewelry because of their Appearance They have a lustrous appearance and come in a variety of colors making them aesthetically pleasing Malleability and ductility They can be shaped into intricate designs and formed into delicate pieces of jewelry Durability Some metals are resistant to scratches and wear ensuring longevity and durability of the jewelry Value Precious metals like gold

and silver are valued for their rarity and beauty making them desirable for jewelry 5 c Give one example of a metal that is used in the manufacture of airplanes and explain why this metal is suitable for this purpose Aluminum is widely used in airplane construction It is chosen because of its Lightness It is a light metal crucial for reducing the weight of the airplane Strength and durability Despite its lightness it is strong and durable enough to withstand the stresses of flight Corrosion resistance Aluminum is resistant to corrosion important for maintaining the structural integrity of the aircraft Remember This guide provides answers and explanations but it is not a replacement for thorough understanding of the concepts Actively engage with the chapter practice additional problems and seek further clarification when needed Good luck with your IGCSE Chemistry studies

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this book is structured to align with the latest syllabus and curriculum guidelines ensuring the content is relevant and rigorous each chapter begins with a clear set of learning objectives providing a roadmap for students to understand what they will achieve by the end of the chapter we have included numerous diagrams illustrations and real life examples to make complex concepts more accessible and engaging

this book was created to help teachers as they instruct students through the master s class chemistry course by master books the teacher is one who guides students through the subject matter helps each student stay on schedule and be organized and is their source of accountability along the way with that in mind this guide provides additional help through the laboratory exercises as well as lessons quizzes and examinations that are provided along with the answers the lessons in this study emphasize working through procedures and problem solving by learning patterns the vocabulary is kept at the essential level practice exercises are given with their answers so that the patterns can be used in problem solving these lessons and laboratory exercises are the result of over 30 years of teaching home school high school students and then working with them as they proceed through college guided labs are provided to enhance instruction of weekly lessons there are many principles and truths given to us in scripture by the god that created the universe and all of the laws by which it functions it is important to see the hand of god and his principles and wisdom as it plays out in chemistry this course integrates what god has told us in the context of this study features each suggested weekly schedule has five easy to manage lessons that combine reading and worksheets worksheets quizzes and tests are perforated and three hole punched materials are easy to tear out hand out grade and store adjust the schedule and materials needed to best work within your educational program space is given for assignments dates there is flexibility in scheduling adapt the days to your school schedule workflow students will read the pages in their book and then complete each section of the teacher guide they should be encouraged to complete as many of the activities and projects as possible as well tests are given at regular intervals with space to record each grade about the author dr dennis englin earned his bachelor s from westmont college his master of science from california state university and his edd from the university of southern california he enjoys teaching animal biology vertebrate biology wildlife biology organismic biology and astronomy at the master s university his

professional memberships include the creation research society the american fisheries association southern california academy of sciences yellowstone association and au sable institute of environmental studies

based on the premise that many if not most reactions in organic chemistry can be explained by variations of fundamental acid base concepts organic chemistry an acid base approach provides a framework for understanding the subject that goes beyond mere memorization the individual steps in many important mechanisms rely on acid base reactions and the ability to see these relationships makes understanding organic chemistry easier using several techniques to develop a relational understanding this textbook helps students fully grasp the essential concepts at the root of organic chemistry providing a practical learning experience with numerous opportunities for self testing the book contains checklists of what students need to know before they begin to study a topic checklists of concepts to be fully understood before moving to the next subject area homework problems directly tied to each concept at the end of each chapter embedded problems with answers throughout the material experimental details and mechanisms for key reactions the reactions and mechanisms contained in the book describe the most fundamental concepts that are used in industry biological chemistry and biochemistry molecular biology and pharmacy the concepts presented constitute the fundamental basis of life processes making them critical to the study of medicine reflecting this emphasis most chapters end with a brief section that describes biological applications for each concept this text provides students with the skills to proceed to the next level of study offering a fundamental understanding of acids and bases applied to organic transformations and organic molecules

nanoparticles exhibit a range of different properties when compared to bulk materials their high surface area to volume ratio makes them particularly attractive for use as catalysts and recent years have seen an explosion of research in this area the ability to fine tune the size and structure of nanoparticles means that it is possible to design catalytic materials for improved activity or specificity as catalysis is one of the key technologies for more sustainable production of both chemicals and energy the past few years have seen increasing numbers of nanomaterials reported for these applications depending on the application a number of different catalyst synthesis and optimization protocols can be used this book provides comprehensive links between the design and fabrication method for nanoparticles and their catalytic performance activity selectivity and stability in various applications presenting an introduction to the concept of catalyst design and recent developments in the preparation and characterisation of nanomaterials followed by several chapters on the design of catalysts for specific applications this book is a valuable resource for researchers working on catalytic reactions industrial processes and nanomaterial applications

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over two decades have passed since the fifth edition of phosphorus chemistry biochemistry and technology major advances in chemistry materials science electronics and medicine have expanded and clarified the role of phosphorus in both our everyday appliances and groundbreaking research significantly expanded updated and reorganized this sixth edition organizes and explains vital phosphorus research and relevant information available in highly specialized reviews and references on select related topics an authoritative and comprehensive review of phosphorus chemistry and related technology phosphorus chemistry biochemistry and technology covers historical academic industrial agricultural military biological and medical aspects of phosphorous furthermore it offers a starting point for more extended studies of the highly specialized branches of phosphorus chemistry although this book deals with a small fraction of the 106 known phosphorus compounds it thoroughly covers the simpler derivatives and most key compounds of economic sociological and biological importance extensively updated and expanded with tables figures equations structural formulae and references it is ideal for scientists in related fields seeking a rapid introduction to phosphorus chemistry

this second edition encyclopedia supplies nearly 350 gold standard articles on the methods practices products and standards influencing the chemical industries it offers expertly written articles on technologies at the forefront of the field to maximize and enhance the research and production phases of current and emerging chemical manufacturing practices and techniques this collecting of information is of vital interest to chemical polymer electrical mechanical and civil engineers as well as chemists and chemical researchers a complete reconceptualization of the classic reference series the encyclopedia of chemical processing and design whose first volume published in 1976 this resource offers extensive a z treatment of the subject in five simultaneously published volumes with comprehensive indexing of all five volumes in the back matter of each tome it includes material on the design of key unit operations involved with chemical processes the design unit operation and integration of reactors and separation systems process system peripherals such as pumps valves and controllers analytical

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this book was designed to serve as a text for lipids low caloric fats and biotechnology have courses in food chemistry in food science pro received a good deal of attention our under grams following the institute of food technolo standing of the functionality of proteins expands gists minimum standards the original idea in with increasing knowledge about their composi the preparation of this book was to present basic tion and structure carbohydrates serve many information on the composition of foods and the functions in foods and the noncaloric dietary chemical and physical characteristics they fiber has assumed an important role undergo during processing storage and han color flavor and texture are important dling the basic principles of food chemistry attributes of food quality and in these areas remain the same but much additional



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fundamentals of forensic chemistry strives to help scientists lawyers students understand how their two disciplines come together for forensic science in the contexts of analytical chemistry related science more generally and the common law systems of canada usa uk the commonwealth in this book forensics is considered more generally

than as only for criminal law workplace health safety and other areas are included and two issues of canadian legal process are argued as essays in the final two chapters

formally established by the epa nearly 15 years ago the concept of green chemistry is beginning to come of age although several books cover green chemistry and chemical engineering none of them transfer green principles to science and technology in general and their impact on the future defining industrial ecology environmental science and tec

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## FAQs

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